UNIVERSITÀ
DEGLI STUDI DI MILANO-BICOCCA

SYLLABUS DEL CORSO

Chimica Fisica

2021-2-E3201Q092

Aims

To provide the basic concepts for the investigation of macroscopic systems, the foresight of spontaneous processes, the definition of equilibrium conditions, the study of chemical kinetics

- ability to learn, stimulated by the constant effort to resume and integrate the previous knowledge of mathematics, physics and chemistry, to analyze complex physical and chemical transformations.

Contents

Thermodynamics: Thermodynamic properties and laws. Applications to: phase transitions; solutions; chemical reactions.

Chemical Kinetics: Rates of chemical reactions. Reaction mechanisms. Activation parameters. Catalysis.

Detailed program

b) Thermodynamics:

- Definition of the main thermodynamic properties (energy, enthalpy, entropy) and references to the laws of thermodynamics.
- Free energy and equilibrium; evaluation of the spontaneity of a process through calculation of the Gibbs free energy variations.
- Phase equilibrium: phase diagrams of pure substances; Clausius Clapeyron equation.
- Mixing equilibrium: ideal and real gas mixtures; mixtures and solutions of ideal and real liquids.
- Chemical equilibrium: variation of Gibbs free energy in reactions; equilibrium constant; van't Hoff equation.

- Rates of chemical reactions. Rate law, rate constant and order of a reaction. Kinetic experiments.
- Methods for determining the reaction order and the rate constant.
- Reaction mechanisms.
- Reaction coordinate, transition state and activation parameters; Arrhenius equation.
- Catalysis; Michaelis Menten equation.

Prerequisites

Prerequisite. Basic knowledge of: general, inorganic and organic Chemistry, Mathematics, Physics.

Propaedeutics. There are no imposed propaedeutics, but it is recommended to have passed the exam of General and Inorganic Chemistry, Mathematics I, and General Physics.

Teaching form

Lessons, 4 credits (32 hours)

Exercises, 2 credits (20 hours)

In the Covid-19 emergency period, lessons and exercises will take place remotely asynchronously with synchronous videoconferencing events.

Textbook and teaching resource

In the e-learning page of the course are provided: the slides presented during the lessons; some exercises with solutions for individual preparation for the written exam; the videotapes of all the lessons and exercises.

Semester

Second semester

Assessment method

The assessment of the Physical Chemistry course consists in a written and an oral examinations. The two examinations can be taken separately.
The written examination consists in the solution of thermodynamic and chemical kinetic exercises and has the objective of verifying the acquired knowledge, the ability to apply such knowledge to the solution of problems, and the ability of judgement. **The oral examination** includes a brief discussion on the written test and an interview on the topics covered in class. It aims to verify the knowledge and understanding of the exam program and the communication skills.The final mark (18-30/30) is obtained by integrating the evaluations of the written and oral examinations.During the Covid-19 emergency period, both tests can be carried out using the WebEx platform and a public link will be provided on the teaching e-learning page for access to the examination of possible virtual spectators.

Office hours

Students reception after e-mail appointment.