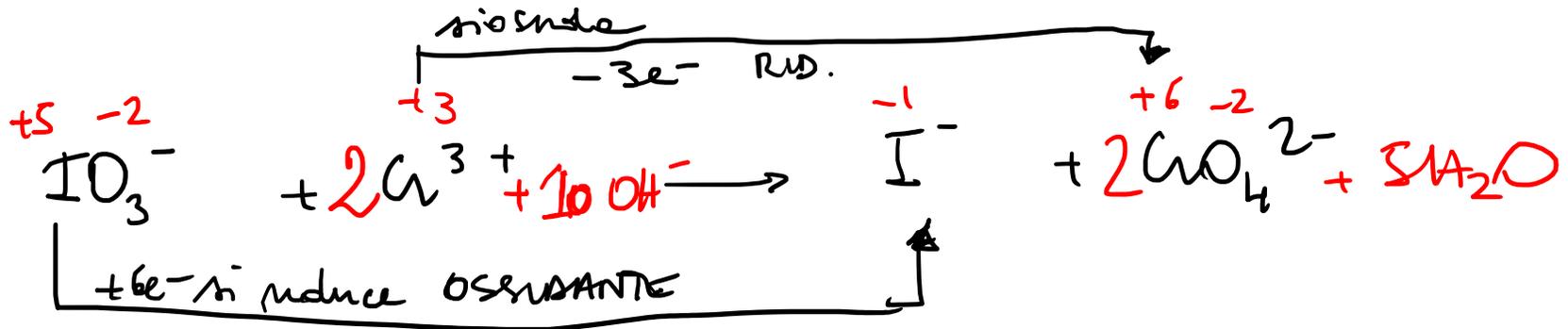
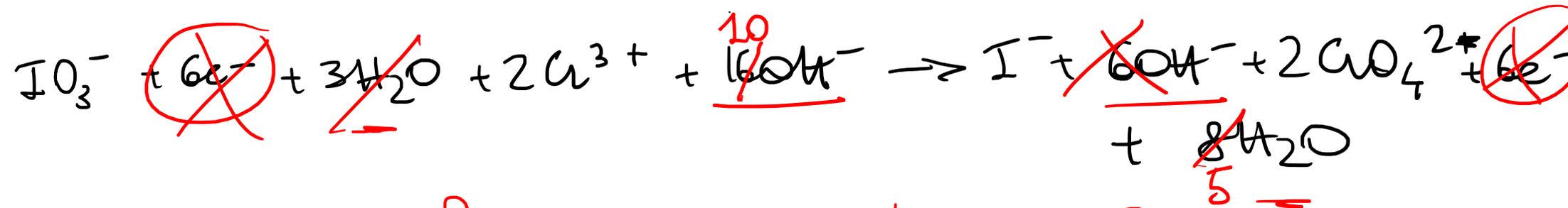
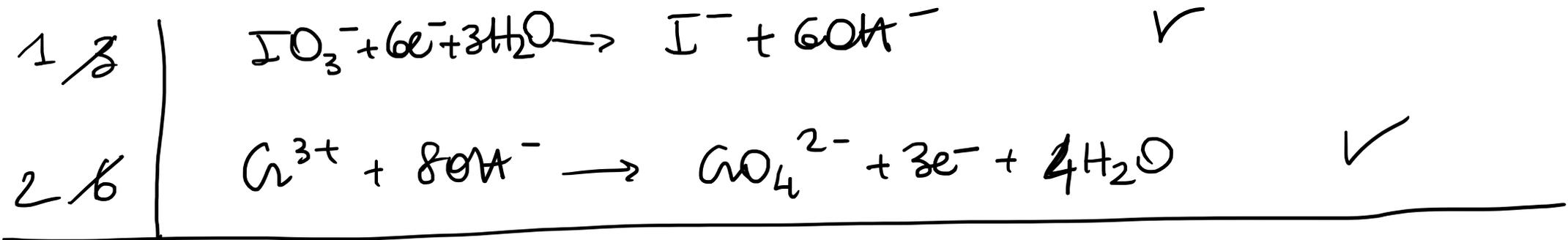


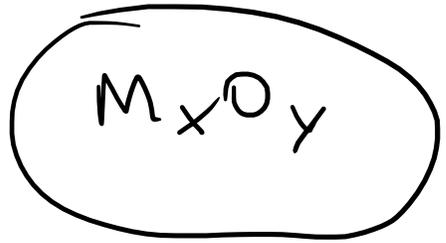
Lezione 7



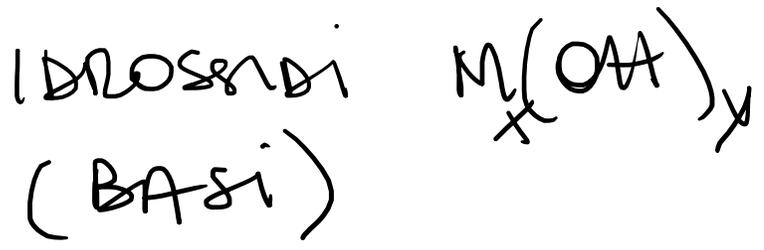
ambiente
basico



Reazione è bilanciata



Composti ionici



→ sorgenti di OH^- in soluzione

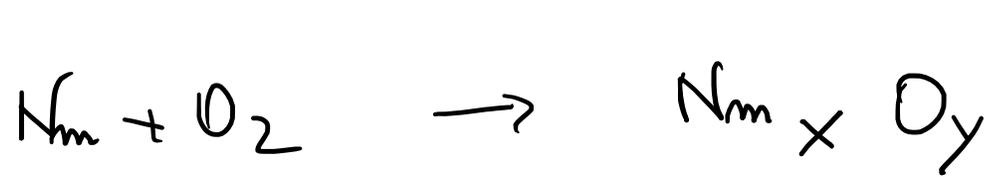
- $NaOH$ $Ca(OH)_2$
- KOH $Mg(OH)_2$
- $LiOH$ $Be(OH)_2$



ossido di CALCIO

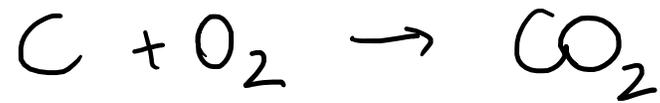
Idrossido di CALCIO



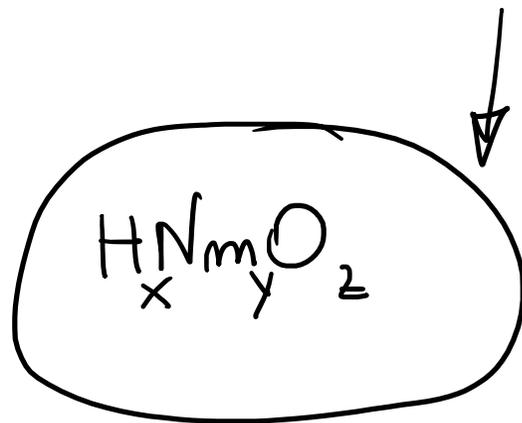
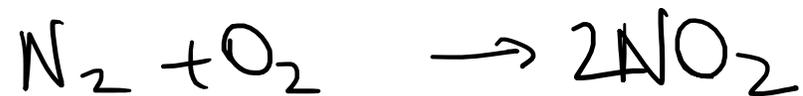


ossidi acidi (ANIDRIDI)

COMPOSTI COVALENTI



DIOSSIDO DI CARBONIO
ANIDRIDE CARBONICA

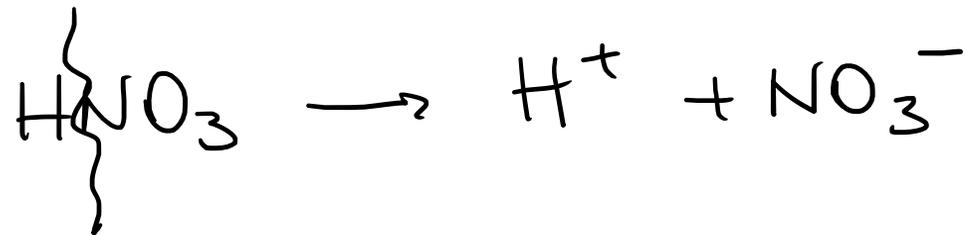


ACIDI (OSERACIDI)
composti covalenti
ma in H_2O

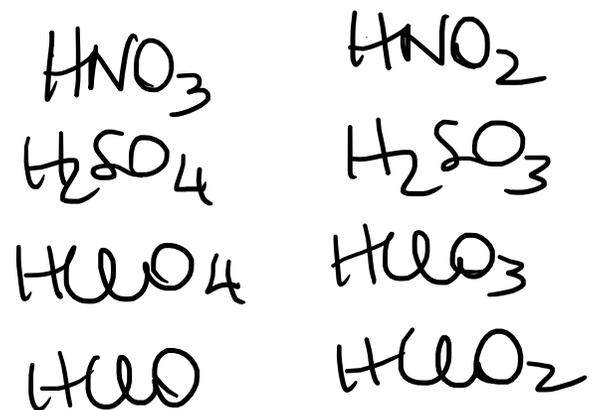
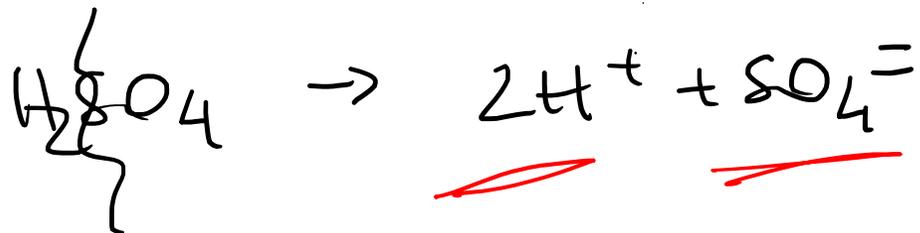
Sorgenti di
 H^+



ACIDO NITRICO

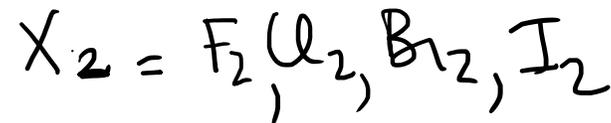


ACIDO SOLFORICO



(7)

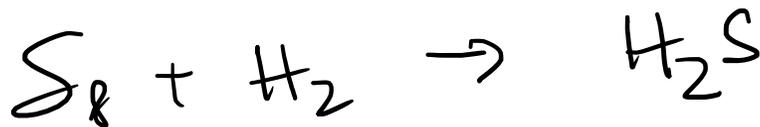
Alogeni (VIIA)



ACIDO CLORIDRICO

↓ sorgente di H^+

HBr HI
HF



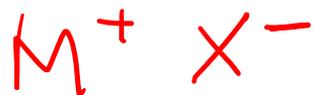
ACIDO SOLFIDRICO



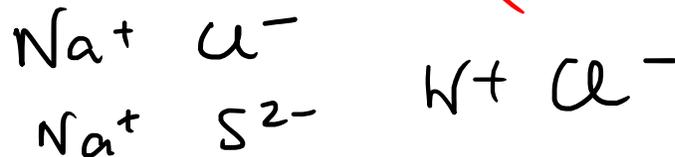
IDRACIDI

SAU

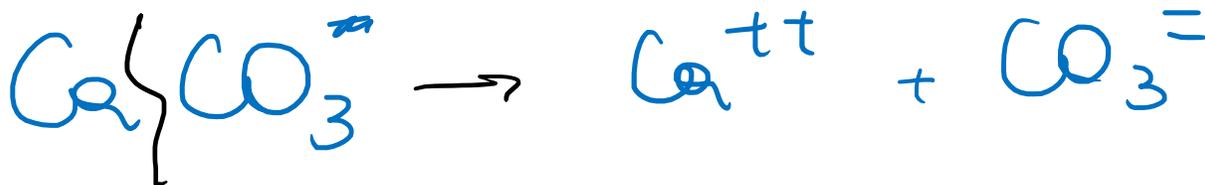
Composti ianici

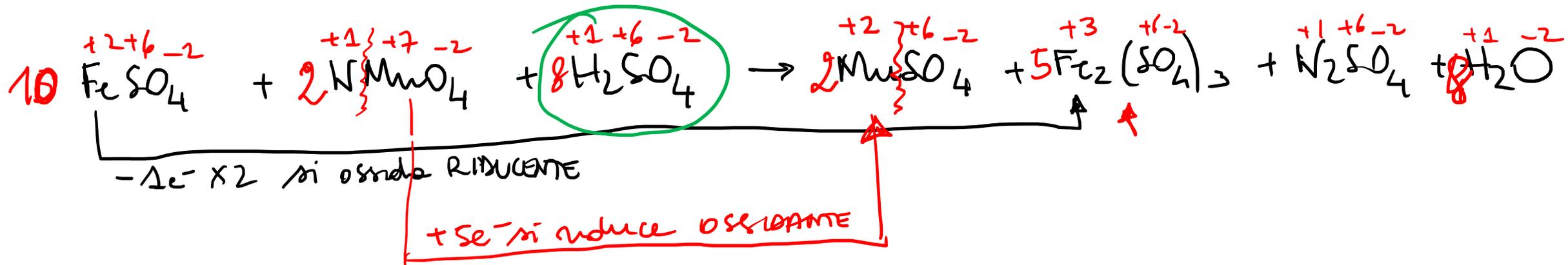


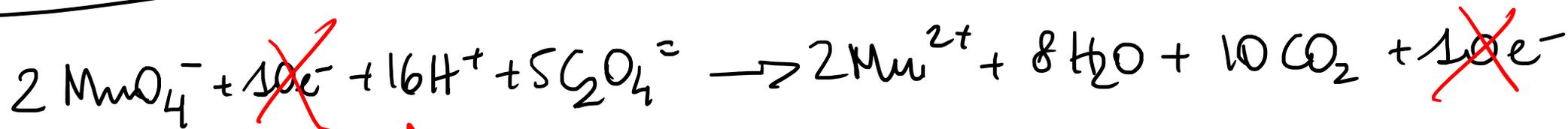
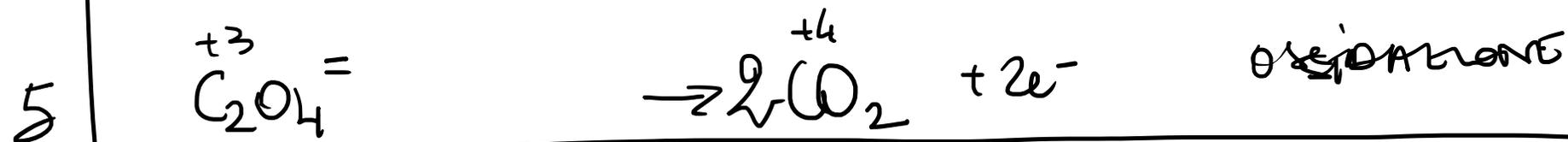
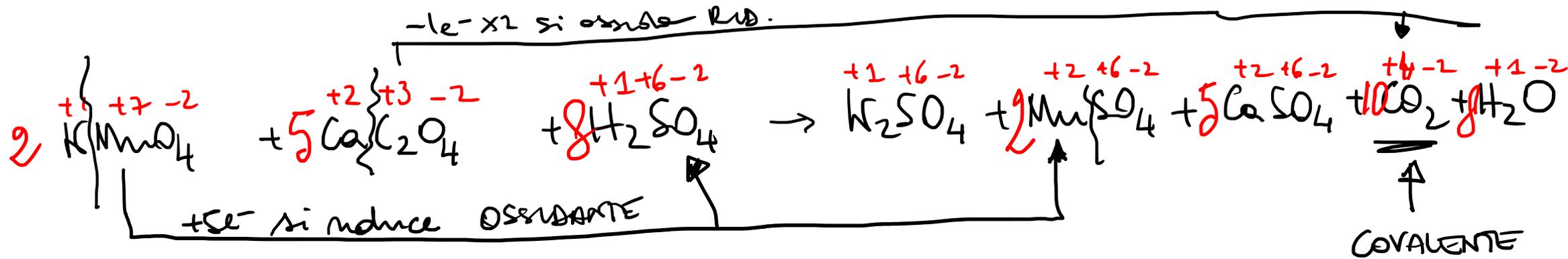
Sale binarice (ioni monoatomici)



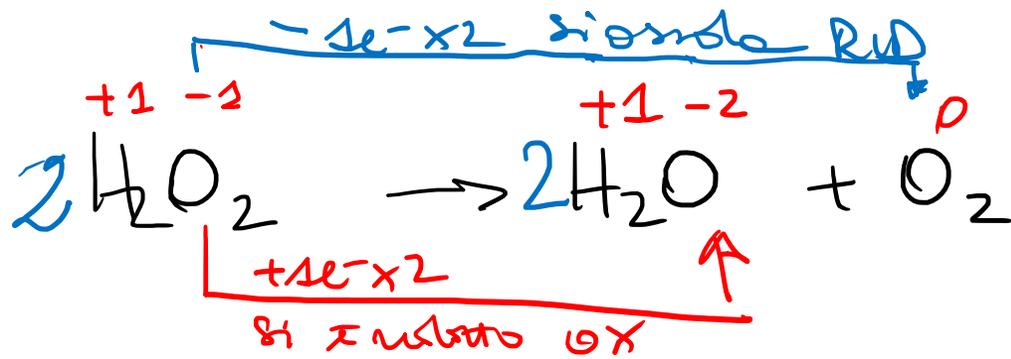
SAU TERNARI (ioni monoatomici + ioni poliatomici)







(M)



H_2O_2
 Acqua ossigenata
 PEROSSIDO DI IDROGENO
 O_2^{2-}

QUANDO UNA STESSA SPECIE
 SI OSSIDA e SI RIDUCE CONTEMPORANEAMENTE
 SI PARLA DI DISPROPORZIONE o DISMUTAZIONE
 o REAZ. DI DISPROPORZIONAMENTO

